

Periodic Table with Chemistry Formulas

sparkcharts

CHEMISTRY FORMULAS

ATOMS AND SUBATOMIC PARTICLES

A. Atomic number: The number of protons in an atom of a given element. In the periodic table, elements are ordered by increasing atomic number.

Ex: ^{12}C $Z =$ atomic number

B. Mass number: The number of protons plus the number of neutrons in an atom of an element, measured in atomic mass units (amu). 1 amu = mass of one proton or neutron = 1.66×10^{-27} kg.

Ex: ^{12}C $12 =$ mass number

Name	Symbol	Mass	Charge
Proton	p^+	1 amu	+1
Neutron	n^0	1 amu	0
Electron	e^-	$1/1837$ amu	-1

C. Isotopes: Various atoms of an element that possess different numbers of neutrons, and therefore different mass numbers.

Ex: Boron has ^{10}B , Boron has ^{11}B

D. Calculating atomic mass: An element's atomic mass number on the periodic table is a weighted average of that element's isotopes.

Ex:

Isotope	Abundance	Mass #	Isotope Mass
Boron-10	20.2%	10	10.0129 amu
Boron-11	79.8%	11	11.0093 amu

Atomic mass = $(20.2\%)(10.0129 \text{ amu}) + (79.8\%)(11.0093 \text{ amu})$

$= 10.81 \text{ amu}$

MOLECULAR MASS

A. Avogadro's number: 6.022×10^{23}

6.022×10^{23} atoms or molecules = 1 mole of that element or substance.

B. The atomic mass number of an element equals the mass (in grams) of one mole of that element.

Ex: One mole of boron has a mass of 10.81 grams.

MOLECULAR FORMULA AND EMPIRICAL FORMULA

A. Molecular formula: The chemical formula of a compound that gives the actual number of atoms of each element in a molecule of the compound.

Ex: The molecular formula for glucose is $\text{C}_6\text{H}_{12}\text{O}_6$.

B. Empirical formula: A "simplified" expression of a chemical formula.

Ex: The empirical formula for glucose is $\text{C}_2\text{H}_4\text{O}$.

BALANCING CHEMICAL EQUATIONS

1. Check for diatomic molecules such as H_2 and O_2 .

2. Balance metals.

3. Balance nonmetals.

4. Balance oxygen.

5. Balance hydrogen.

6. You may change only coefficients, never the formula subscripts! Recheck all atoms to be sure you have equal numbers on both sides of the equation.

7. Rewrite the equation so that coefficients are the smallest whole numbers possible.

PERCENTAGE AND PPM

A. Percentage composition: Tells how much mass any object or element contributes to a compound.

Percentage composition = $\frac{\text{mass of element}}{\text{mass of compound}} \times 100$

B. Parts per million (ppm): In a homogeneous solution, the number of grams of a

particular compound that would be present in a one million gram sample of the compound or solution.

TEMPERATURE CONVERSION

A. Converting between Fahrenheit and Celsius

$$\text{Fahrenheit} = \frac{9}{5}(\text{Celsius}) + 32^\circ\text{F}$$

$$\text{Celsius} = \frac{5}{9}(\text{Fahrenheit} - 32^\circ\text{F})$$

B. Kelvin: The temperature in K, absolute zero, is the theoretical lowest possible temperature, at which all molecular activity ceases.

$$0 \text{ K} = -273.15^\circ\text{C} = -459.67^\circ\text{F}$$

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$$\text{C}_p = \frac{Q}{m \Delta T}$$

where Q is the amount the heating point is lowered, k_p is a proportionality constant specific to the solvent, and m is the mass of the solution.

C. Boiling point elevation:

$$\Delta T_b = k_b m$$

ΔT_b is the amount the boiling point is elevated, k_b is a proportionality constant specific to the solvent, and m is the molality of the solution.

D. Osmotic pressure:

$$\pi = iMRT$$

where π is the osmotic pressure, M is the molarity of the solution, R is the ideal gas constant, and T is the temperature in Kelvin.

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concentration with respect to A , denoting the reaction rate by independent of $[A]$.

Percentage yield = $\frac{\text{actual yield}}{\text{theoretical yield}} \times 100$

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Percentage yield = $\frac{\text{actual yield}}{\text{theoretical yield}} \times 10$

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