

## Chemical Periodicity

Name Mitchell Wade Section Number 015

**OBJECTIVE:** To demonstrate the periodic law.

**CONCEPTS:** Periodic law, Periodic trends

**INTRODUCTION:** In the 1860's about 60 elements were known. From their physical properties, the elements were classified as metals, nonmetals, or metalloids. Metals have low electrical resistance (high conductivity), have a metallic luster, and are malleable. Non-metals have high resistance, low luster, and are brittle. Metalloids have properties between those of non-metals and metals. Chemists of the time further classified the elements into groups. The elements in a particular group have similar chemical and/or physical properties. This classification led to the discovery of the periodic table and the periodic law. In the periodic table, elements are assigned to a period and a group—groups are the columns and the periods are the rows of the table. The periodic law states that the properties of the chemical elements vary with the atomic number in a systematic way. We now understand that the properties depend on the electronic structures of the atoms.

### PROCEDURE:

1. Find the station labeled "Halogens: states of matter." Observe the states of matter of the three samples. How does the state of matter change as you go down the column of the periodic table?

Starts as gas and goes to solid

The melting points of the halogen elements that you observed are listed below. Based on the trend that you see, estimate the melting point of Fluorine.

| Element | Melting Point ( $^{\circ}\text{C}$ ) |
|---------|--------------------------------------|
| F       | -200                                 |
| Cl      | -101.5                               |
| Br      | -7.35                                |
| I       | 113.75                               |

2. Find the station labeled "Metallic Solutions." Fill out the chart below. For the unlabeled mystery elements, you will have to guess which type of metal they are (transition, alkali, or alkaline earth) based on the trends you see.

| Solution Number | Element | Color of Water Solution | Type (transition metal, alkali metal, alkaline earth metal) |
|-----------------|---------|-------------------------|-------------------------------------------------------------|
| 1               | Co      | Red                     | Transition metal                                            |
| 2               | Ba      | Clear                   | Alkaline earth metal                                        |
| 3               | K       | clear                   | Alkali metal                                                |
| 4               | Ni      | green                   | Transition metal                                            |
| 5               | Na      | clear                   | Alkali metal                                                |
| 6               | Cu      | blue                    | Transition metal                                            |
| 7               | Mg      | clear                   | Alkaline earth metal                                        |

# Periodicity Lab Report

**L Towne**

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located on the right side of the light blue bar.

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