

PERIODIC TABLE BASIC

Reading the Periodic Table:

Objectives:

- Explain why elements in a group have similar properties.
 - identify the Ksur blocks of the percodic lable.

The Periodic Table:

Our current periodic table tets 109 erements. It is arranged ... economy to the periodic law:

- Elements are arranged in order of increasing atomic number, their physical and chemical properties show a periodic pattern.
- Elements that have similar properties are arranged in proups or families - vertical columns.
- The horizontal nows are called periods.
- Each period contains more elements than the previous one. This has to do with the increasing energy levels. (n=1, ...) of the atom.
- podod 5 2 elements.
- period 2 & 2 6 elements
- period 4 5 5 18 elements
- period 6 32 elements

Labeling and Naming:

There are currently 3 numbering solumnes used in chemistry, for the ordering of groups.

- Bromer M VIIIA & 10-VIIIB
- Avidic 1A-2A, 38-88, 18-28, 3A-6A
- w Neuropher 5-58

Families:

- 1A adkasii mensisis
- 2A alkalinet earth metate.
- ZA hallogeme
- EA noble gaves
- 4A sometimes known as the Carbon group

Exception - Hydrogen can be beder placed in group ZA, Athough, some books place it in both.

Metarro & Nicon-metalis:

Metals include the areas shaded light blue and include the families SA.5.2A. They share mony characteristics.

- Rystler (whome)
- good conductors of heat and electricity.
- Sybically solid at room temperature.
- Inost are malleader and ductile.

Non-metals are shaded light red in rocal periodic tables and form a diagonal form C-6 to Rn-86 and include all the elements of that conner including H. They share some of these characteristics.

- ido ricot posisiena kunder
- poor conductors of heat and electricity
- neither maleuple or ductile
- many are passes at room temperature.
- man be sold, semisolid, gas, or liquid (fir). One is the hardest material on Earth

Semimentals or metallisets are shaded light light purple in most periodic tables form a diagonal from \$-5 to At-65 and include all the elements Ge-32 and St-51. They share some of these characteristics.

have some properties of metals and some of non-metals

Electron Configuration:

As you learned series, the number of protons in an atom is equal to the number of electrons in the atom. If it is neutral. The periodic table allows you to predict the bonding of an atom by allowing you to see the number of electrons in the valence orbital. The valence orbital is the highest, numbers orbital of the atom and is the only one involved in the bonding of the atom. Here are some hints for predicting satisface electrons.

- Elements of the same group (column) have the same visionce electrons or configuration.
- s-block is group SA.& 2A.
- p-block in-group 3A to 6A.
- These are the impresentative elements or main group standards.
- id-block is group 38 to 28 (look at the periodic table above because the numbers are not sequential)
- These are the transition metals.
- 5-block is Societal below the main table
 - These are the inner transition metals or rare earth stamperts.

To same space on electron configuration notation you can use the abbreviated notation by listing the inner gas core and then this velocie electrons.

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