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Coupon Code Wifi 7 Router Last 90 Days

LL Leslie



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web solution 1 the two half reactions balanced as if in acidic solution $2\text{CN}^- + \text{Au} \rightarrow 2\text{e}^- + 2\text{H}_2\text{O} + 2\text{AuCN}$ make electrons equal convert to basic solution $4\text{CN}^- + 2\text{Au} \rightarrow 2\text{e}^- + 2\text{H}_2\text{O} + 2\text{AuCN}$ multiplied by a factor of 2 $2\text{e}^- + 2\text{H}_2\text{O} + 2\text{AuCN} \rightarrow 2\text{OH}^- + 2\text{H}_2\text{O} + 2\text{AuCN}$ the final answer

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bromate v ions

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