

Be Prepared

FOR THE DIGITAL SAT

A WALKTHROUGH OF THE DIGITAL SAT INTERFACE, INCLUDING:

1

The Desmos Calculator, Answer Elimination Tool, and Annotation Feature

2

An Adaptive SAT Practice Test in the book simulates the section-adaptive format of the Digital SAT

3

Strategies on all question types like Evidence, Words-in-Context, Rhetorical Synthesis, Problem Solving and Data Analysis, and Geometry and Trigonometry

4

Hundreds of practice drills representing Reading, Writing, and Math sections on the Digital SAT

PRACTICE TEST 2: ADAPTIVE

160 DIGITAL SAT STUDY GUIDE

Adaptive Test Overview

As we about made on Chapter 10's Practice Test, this will describe the section-adaptive format of the Digital SAT. Depending on your performance on the first module of each section, you will receive a more difficult or less difficult second module of each section. (To select a module, you'll see a potential question selected. Once we're overlapping questions in any of the modules—i.e., from the Reading and Writing Module 20, for example, you may not go back and to Reading and Writing Module 24 for a new question set.) The format of the test is outlined in the table below. The sections are outlined throughout the book sections. Good luck!

SAT Section	Format
Reading and Writing Module 1	20 Questions, 20 Questions, Standard Difficulty Based on module after completion to determine whether you'll go to 24 or 28.
Reading and Writing Module 24 or 28	20 Questions, 20 Questions, Adaptive Difficulty If you get 17 or fewer questions correct on the first module, go to the less-challenging Module 24. If you get 18 or more questions correct on the first module, go to the more-challenging Module 28.
Math Module 1	
Math Module 1	20 Questions, 20 Questions, Standard Difficulty Based on module after completion to determine whether you'll go to 24 or 28.
Math Module 24 or 28	20 Questions, 20 Questions, Adaptive Difficulty If you get 17 or fewer questions correct on the first module, go to the less-challenging Module 24. If you get 18 or more questions correct on the first module, go to the more-challenging Module 28.

Summary

This chapter introduces the Digital SAT practice test, which is the first of the two practice tests in the book. It is the first of the two practice tests in the book.

Sat Practice Guide Warranty

Annelies Wilder-Smith

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the author's name.

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 $2 \text{ CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g}) + \text{C}_2\text{H}_2(\text{g}) \rightarrow \text{O}_2(\text{g}) + \text{C}_2\text{H}_6(\text{g})$ $\Delta h = 94.5 \text{ kJ}$
 $\text{H}_2\text{O}(\text{g}) + \text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l})$ $\Delta h = 71.2 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + 3 \text{H}_2\text{O}(\text{g})$ $\Delta h = 283 \text{ kJ}$

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 $\text{C}_2\text{H}_4(\text{g}) + 3 \text{O}_2(\text{g}) \rightarrow 2 \text{CO}_2(\text{g}) + 2 \text{H}_2\text{O}(\text{l})$ $\Delta h = 1411 \text{ kJ}$ mole
 $\text{C}_2\text{H}_6(\text{g}) + 7 \text{O}_2(\text{g}) \rightarrow 2 \text{CO}_2(\text{g}) + 3 \text{H}_2\text{O}(\text{l})$ $\Delta h = 1560 \text{ kJ}$ mole
 $\text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l})$ $\Delta h = 285 \text{ kJ}$ mole

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 hess s law example 12.1 determine Δh for the target reaction
 $\text{C}_2\text{H}_4(\text{g}) + \text{H}_2(\text{g}) \rightarrow \text{C}_2\text{H}_6(\text{g})$
 given the following information

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 hess s law is written as $\Delta h_{\text{target}} = \Delta h_1 + \Delta h_2 + \Delta h_3$ or $\Delta h_{\text{target}} = \sum \Delta h_{\text{known}}$

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 $\text{Mg}(\text{s}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{MgO}(\text{s})$ we used

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 cohen 2016 the equations for the reactions in the experiment done are as follows
 $\text{NaOH}(\text{s}) + \text{Na}(\text{aq}) \rightarrow \text{OH}(\text{aq})$

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web thus for reaction 9 2 in which 1 mole of H_3O^+ and 1 mole of OH^- combine to form 2 moles of H_2O $\Delta h = 55.8 \text{ kJ}$ as shown note that the enthalpy change $\Delta h_{\text{products}} - \Delta h_{\text{reactants}}$ is positive if heat is absorbed that is if $\Delta h_{\text{products}}$

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