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Purchase Date: March 2024

Serial Number: [REDACTED]

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convert to basic solution $4\text{CN}^- + 2\text{Au} + \text{CN}^- + 2\text{e}^-$ multiplied by a factor of 2 $2\text{e}^- + 2\text{H}_2\text{O} \rightarrow \text{O}_2 + 2\text{OH}^-$ 3 the final answer

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation Ni(s) $\rightarrow \text{Ni}^{2+} + 2\text{e}^-$ now that the oxidation half reaction is balanced it tells us that two electrons are produced for

every atom of nickel oxidized but where do those electrons go

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web balancing redox reactions in acidic solution problems 1 10 problem 1 $\text{Cr}_2\text{O}_7^{2-}$ Fe^{2+} Cr^{3+} Fe^{3+} solution 1 balanced half reactions 6e^- 14H^+ $\text{Cr}_2\text{O}_7^{2-}$ 2Cr^{3+} $7\text{H}_2\text{O}$ Fe^{2+} Fe^{3+} e^- 2 equalize the electrons 6e^- 14H^+ $\text{Cr}_2\text{O}_7^{2-}$ 2Cr^{3+} $7\text{H}_2\text{O}$ 6Fe^{2+} 6Fe^{3+} 6e^- multiplied by a factor of 6

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web in the redox equation below that occurs in our stomach which of the following pairs identifies the reducing and oxidizing agents respectively $\text{C}_6\text{H}_8\text{O}_6$ $2\text{H}_2\text{NO}_2$ $\text{C}_6\text{H}_6\text{O}_6$ $2\text{H}_2\text{O}$ 2NO choose 1 answer $\text{C}_6\text{H}_8\text{O}_6$ and NO_2 a $\text{C}_6\text{H}_8\text{O}_6$ and NO_2 no 2 and $\text{C}_6\text{H}_8\text{O}_6$ b NO_2 and $\text{C}_6\text{H}_8\text{O}_6$ h and NO_2 c h and

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