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temperature change plugging in the values given in the problem you get $q_{\text{water}} = 418 \text{ J/g} \times 110 \text{ g} \times 26.6 \text{ }^{\circ}\text{C} = 11800 \text{ J}$
 $550 \text{ J} = \Delta h \times q_{\text{water}} = 550 \text{ J}$

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 for water $4184 \text{ J/g} \times 1$ what is the specific heat of aluminum if the temperature of a 28.4 g sample of aluminum is increased
 by $8.1 \text{ }^{\circ}\text{C}$ when 207 J of heat is added 2

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 $147314 \Delta H_{\text{rxn}} = q_{\text{rxn}} / n$ $q_{\text{calorimeter}} = mcs\Delta T$

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web use the information in this chart to answer the following questions 1 if i have 125 grams of water at a temperature of
 $25.0 \text{ }^{\circ}\text{C}$ and heat it to $145.0 \text{ }^{\circ}\text{C}$ what will the change in enthalpy be 2 sketch the heating curve temperature vs time for the
 process in 1 3 i have 18.7 grams of an unknown liquid chemical at a temperature of $35.0 \text{ }^{\circ}\text{C}$

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 of 1.00 M NaOH at $25.0 \text{ }^{\circ}\text{C}$ in a foam cup calorimeter a reaction occurs and the temperature rises to $32.0 \text{ }^{\circ}\text{C}$ calculate the
 enthalpy change in kJ for this reaction assume the densities of the solutions are 1.00 g/ml and that

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