

Oxidation and Reduction

Reactivity of metals with oxygen.

Reactivity of metals with oxygen			
- Conclude by the flame or glow produced when metal is heated with oxygen. - More reactive metal with oxygen, more brightly and rapidly and vigorously the metal burns.			
Metal	Intensity of flame / glow	Observation	
		Colour of hot oxide	Colour of cold oxide
Magnesium	Burns very rapidly Very bright white flame produced. $2 \text{Mg} + \text{O}_2 \rightarrow 2 \text{MgO}$	White	White
Zinc	Burns rapidly Bright glow $2 \text{Zn} + \text{O}_2 \rightarrow 2 \text{ZnO}$	Yellow	White
Iron	Burns less rapidly The glow is less bright than the burning zinc. $4 \text{Fe} + 3 \text{O}_2 \rightarrow 2 \text{Fe}_2\text{O}_3$	Reddish-brown	Reddish-brown
Lead	Burns slowly Faint glow $2 \text{Pb} + \text{O}_2 \rightarrow 2 \text{PbO}$	Brown	Yellow
Copper	Faint glow $2 \text{Cu} + \text{O}_2 \rightarrow 2 \text{CuO}$	Black	Black
For Experiment Oxygen is supplied by heating - potassium manganate (VII) $2\text{KMnO}_4 \xrightarrow{\text{heat}} \text{K}_2\text{MnO}_4 + \text{MnO}_2 + \text{O}_2$ - potassium nitrate $2\text{KNO}_3 \xrightarrow{\text{heat}} 2\text{KNO}_2 + \text{O}_2$ - mixture of potassium chlorate (V) and manganese (IV) oxide $2\text{KClO}_3 + \xrightarrow[\text{heat}]{\text{MnO}_2} 2\text{KCl} + 3\text{O}_2$			
Glass wool is used to prevent potassium manganate (VII) from spilling over metal powder. If potassium manganate (VII) is mixed with metal powder, an explosion may occur during heating.			

Oxidation Reduction Guide

JA Banks



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