

Answers: Oxidation Number

1. N in N_2O_3 N^{3+}
2. S in H_2SO_4 S^{6+}
3. C C^0
4. C in CO C^{2+}
5. Na in NaCl Na^+
6. H in H_2O H^+
7. Ba in BaCl_2 Ba^{2+}
8. N in NO_2^- N^{3+}
9. S in Al_2S_3 S^{2-}
10. S in HSO_4^- S^{6+}
11. Cl in $\text{Fe}(\text{ClO}_2)_3$ Cl^{3+}
12. Fe in $\text{Fe}(\text{ClO}_2)_3$ Fe^{3+}
13. N in NO_3^- N^{5+}
14. Cu^{2+} Cu^{2+}
15. Zn^{2+} Zn^{2+}
16. C in CH_4 C^{4+}
17. Mn in MnO_2 Mn^{4+}
18. S in SO_3^{2-} S^{4+}
19. Mg^{2+} Mg^{2+}
20. Cl^- Cl^-
21. O_2 O_2^0
22. P_4 P_4^0
23. Na in Na_2S Na^+
24. S in H_2S S^{2-}
25. Ca^{2+} Ca^{2+}
26. C in CN^- C^{4+}
27. H in OH^- H^+
28. Mn in KMnO_4 Mn^{7+}
29. I in $\text{Mg}(\text{IO}_3)_2$ I^{5+}
30. C in $\text{C}_2\text{O}_4^{2-}$ C^{2+}

Oxidation Number Practice Answer Key

Guillaume Favre

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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