Reduction-Oxidation Reactions Worksheet

Part I: Identifying parts of Simple Red-ox reactions

For the equations below, all in aqueous solution:

- 1) assign oxidation numbers to each atom
- identify the reactant oxidized,
- 3) the reactant reduced,
- the oxidizing agent,
- 5) the reducing agent,
- write the balanced oxidation and reduction half reactions.

Example: oxidized reduced

Mg° + Br₂° → Mg*² Br¹²

Reducing agent oxidizing agent

Oxidation half reaction: Mg° → Mg*² + 2 e°

Reduction half reaction: 2 e° + Br₂° → 2Br°

Fe + Zn⁺² → Fe⁺² + Zn

Al + Fe⁺² → Al⁺³ + Fe

3. Cu + AgNO₃ \rightarrow Cu(NO₃)₂ + Ag

SiCl₄ (I) + Mg (s) → MgCl₂ (s) + Si (s)

Oxidation Reduction Reactions Study Guide Answers

Neil E. Schore

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