

Name: \_\_\_\_\_

## Atomic Structure & the Periodic Table Study Guide

### Atomic Structure

| Particle | Charge             | Where are they found?                        |
|----------|--------------------|----------------------------------------------|
| Proton   | Positive           | Nucleus                                      |
| Neutron  | Neutral/ no charge | Nucleus                                      |
| Electron | negative           | Electron cloud/energy levels outside nucleus |

1. What is the rule for writing chemical symbols like Na or Ca? **Capitalize the first letter; lower case second letter**
2. What are the vertical columns on the periodic table called? Groups or families
  - a. How many of these are on the periodic table? 18
  - b. Why are the elements in these columns? They have the same chemical properties (& same # of valence electrons)
  - c. Elements in the same group have the same number of valence electrons.
3. What are the horizontal rows on the periodic table called? periods
  - a. How many of these are on the periodic table? 7
  - b. Elements in these rows have the same number of energy (*electron shells*) levels.
4. What are valence electrons? **Electrons in the outer energy level**
5. How many electrons are in each energy level?  
1<sup>st</sup> energy level 2      2<sup>nd</sup> energy level 8      3<sup>rd</sup> energy level 8
6. The atomic number is equal to the # of protons and also # of electrons
7. The atomic mass is equal to the # protons + # of neutrons
8. How do you find the number of neutrons? Atomic mass – atomic number
9. We view the order of the periodic table today according to the atomic number

# Periodic Table Study Guide

**David E. Sadava, Gordon H.  
Orians, Craig Heller, William K. Purves**



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