

BEGINNER PILATES

A woman with long blonde hair, wearing a white tank top and white leggings, is lying on her back on a yellow mat. Her legs are raised and bent at a 90-degree angle, with her feet flat against a white wall. Her arms are extended to the sides on the floor. The background shows a bright room with a large window and a potted plant.

10 MIN

WALL PILATES

Yoga For Beginners Pilates At Home Same Day Delivery

Kate J. Chase



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 $2 \text{ CO}_2 \text{ g} + \text{H}_2\text{O g} + \text{C}_2\text{H}_2 \text{ g} \rightarrow 2 \text{ CO}_2 \text{ g} + 2 \text{ H}_2 \text{ g} + \text{C}_2\text{H}_6 \text{ g}$ Δh 94 5 kJ
 $\text{H}_2\text{O g} + \text{H}_2 \text{ g} + \frac{1}{2} \text{O}_2 \text{ g} \rightarrow \text{H}_2\text{O g}$ Δh 71 2 kJ
 $\text{C}_2\text{H}_6 \text{ g} + \frac{7}{2} \text{O}_2 \text{ g} \rightarrow 2 \text{ CO}_2 \text{ g} + 3 \text{ H}_2\text{O g}$ Δh 283 kJ
g 1 $\text{C}_2\text{H}_6 \text{ g}$

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 $\text{C}_2\text{H}_4 \text{ g} + 3 \text{ O}_2 \text{ g} \rightarrow 2 \text{ CO}_2 \text{ g} + 2 \text{ H}_2\text{O l}$ Δh 1411 kJ
mole $\text{C}_2\text{H}_6 \text{ g} + 7 \text{ O}_2 \text{ g} \rightarrow 2 \text{ CO}_2 \text{ g} + 3 \text{ H}_2\text{O l}$ Δh 1560 kJ
mole $\text{H}_2 \text{ g} + \frac{1}{2} \text{O}_2 \text{ g} \rightarrow \text{H}_2\text{O l}$ Δh 285 kJ
mole

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web thus for reaction 9 2 in which 1 mole of H_3O^+ and 1 mole of OH^- combine to form 2 moles of H_2O $\Delta H = 55.8 \text{ kJ}$ as shown note that the enthalpy change $\Delta H_{\text{products}} - \Delta H_{\text{reactants}}$ is positive if heat is absorbed that is if $\Delta H_{\text{products}} > \Delta H_{\text{reactants}}$

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web objective the purpose of the lab is to combine two reactants in the calorimeter and measure the heats of reaction in solution eventually allowing an individual to prove or disprove hess s law introduction thermodynamics is the study of energy exchanged between the system and the surroundings

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web 3 12 hess s law in this experiment you will measure the amount of heat released in these three related exothermic reactions $\text{NaOH}(\text{s}) + \text{Na}^+(\text{aq}) + \text{OH}^-(\text{aq}) \rightarrow \text{H}_2\text{O}(\text{l})$ $2\text{NaOH}(\text{s}) + \text{H}^+(\text{aq}) + \text{Cl}^-(\text{aq}) \rightarrow \text{H}_2\text{O}(\text{l})$ $\text{Na}^+(\text{aq}) + \text{Cl}^-(\text{aq}) \rightarrow \text{H}_2\text{O}(\text{l})$

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