

## 17: Molecular Geometry - VSEPR

### Key Structure Terms

- Valence Shell:** Electrons in the outermost shell that are involved in bonding.
- Lewis Structure:** A 2D representation of a molecule and its bonds.
- Lone Pair:** Pair of electrons not being shared in a bond.
- Bonding Pair:** Pair of electrons used to make a bond. Both atoms sharing the electrons can "count" them in their valence shell.
- Single bond:** One pair of shared electrons.
- Double bond:** Two pairs of shared electrons. Shorter and stronger than a single bond.
- Triple bond:** Three pairs of shared electrons. Shorter and stronger than a double bond.
- Valence Shell Electron Pair Repulsion Theory (VSEPR):** Bonds and lone pairs are electrons, all electrons are negative, negative things repel other negative things, therefore all bonds and lone pairs arrange themselves in 3D as far away from each other as possible.
- Electron Geometry:** 3D structure of a molecule determined by counting the electron regions around a central atom (bonds and lone pairs).
- Electron Region:** Each bond (single, double or triple) and lone pair count as "1" electron region.
- Molecular Geometry:** 3D structure determined by the atoms bonded to the central atom.
- Ligand:** Atoms bonded to the central atom.

### VSEPR Theory and Geometry

**Electron geometry** is determined by looking at the number of electron regions around the central atom.

**Molecular geometry** is determined by looking at the number of atoms bonded to the central atom (ligand) and the number of lone pairs around the central atom.

**Electronic Geometry Mnemonic:** Linear(2) – Trigonal Planar(3) – Tetrahedral(4) – Trigonal Bipyramidal(5) – Octahedral(6) = "Long DIP To TRiBe QVereases."

**Examples:**



Electron geometry: Tetrahedron  
Molecular geometry: Tetrahedron



Electron geometry: Tetrahedron  
Molecular geometry: Trigonal pyramidal

### Molecular Geometries

A = central atom; X = ligands; E = lone pairs

| Electron Regions | Molecular Formula                                                       | Name                 | Shape |
|------------------|-------------------------------------------------------------------------|----------------------|-------|
| 2                | $AX_2$<br>( $BeCl_2$ , $CO_2$ )                                         | Linear               |       |
| 3                | $AX_3$<br>( $BF_3$ , $CO_3^{2-}$ , $NO_3^-$ , $SO_3$ )                  | Trigonal Planar      |       |
|                  | $AX_2E$<br>( $NO_2^+$ , $SO_2$ , $O_3$ )                                | Bent                 |       |
| 4                | $AX_4$<br>( $CH_4$ , $NH_4^+$ , $PO_4^{3-}$ , $SO_4^{2-}$ , $ClO_4^-$ ) | Tetrahedron          |       |
|                  | $AX_3E$<br>( $NH_3$ , $H_3O^+$ , $PCl_3$ , $SO_3^{2-}$ )                | Trigonal pyramidal   |       |
|                  | $AX_2E_2$<br>( $H_2O$ , $ClO_2$ , $OF_2$ , $SCl_2$ )                    | Bent                 |       |
| 5                | $AX_5$<br>( $PCl_5$ )                                                   | Trigonal bipyramidal |       |
|                  | $AX_4E$<br>( $SF_4$ , $SCl_5$ )                                         | See-saw              |       |
|                  | $AX_3E_2$<br>( $ClF_3$ , $ICl_3$ )                                      | T-shaped             |       |
|                  | $AX_2E_3$<br>( $XeF_2$ , $I_2$ )                                        | Linear               |       |
| 6                | $AX_6$<br>( $SF_6$ , $PCl_6^-$ )                                        | Octahedron           |       |
|                  | $AX_5E$<br>( $BrF_5$ , $IF_5$ )                                         | Square pyramidal     |       |
|                  | $AX_4E_2$<br>( $XeF_4$ )                                                | Square planar        |       |

**How to Use This Cheat Sheet:** These are the keys related to this topic. Try to read through it carefully twice then write it out on a blank sheet of paper. Review it again before the exams.

# Molecular Structure And Bonding Cheat Sheet

**Kaplan Test Prep**

A decorative graphic consisting of a light blue horizontal bar with a rounded right end, and a red circular shape partially visible behind it.

## **Molecular Structure And Bonding Cheat Sheet:**

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