

POLARITY REVIEW

1. Molarity is a type of _____, similar to g/L , mg/mL , cm^3/cm^3 , etc. etc. etc.
2. Molarity is defined as M of _____ dissolved within one _____ of solution.
3. M is the unit for molarity, but the unit for molarity can also be written as:
4. Analyze the images below and identify the molarity of the solutions pictured.

2 moles NaCl dissolved in 2 liters of solution



_____ M

4.0 moles NaCl dissolved in 1.000 L of solution



_____ M

4 moles NaCl dissolved in 0.5 L of solution



_____ M

0.5 moles NaCl dissolved in 0.5 liters of solution



_____ M

2 moles NaCl dissolved in 0.25 L of solution



_____ M

1 mole NaCl dissolved in 1.000 L of solution



_____ M

5. Analyze the images below and identify how many total moles of the compound are present

1 L of 0.500 solution



_____ moles

1000 mL of 1.000 solution



_____ moles

500 mL of 1.000 solution



_____ moles

0.5 L of 1.000 solution



_____ moles

2000 mL of 0.500 solution



_____ moles

0.25 L of 1.000 solution



_____ moles

6. Analyze the images below and answer the prompts accordingly. Be sure to include the correct units and identity of the solution on question. Ex: 2 moles CH_4 or 3M MgO

How many moles of NaCl are in 1000 mL of a 0.500 solution?



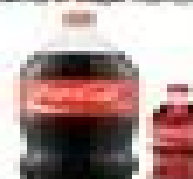
What is the concentration in molarity of 2000 mL of a solution containing 10 moles of NaCl ?



How many moles of NaCl are in 1 L of a 0.500 solution?



Answer the below 3a and 3b with the number of grams. Ex: 100. The total mass (weight) of the solution of Coca Cola is about 100 g. What is a percent of sugar in 100 grams? What is the density of sugar NaCl in this drink?



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