

Molarity Calculations

Calculate the molarities of the following solutions:



- 1) 2.3 moles of sodium chloride in 0.45 liters of solution.

$$M = \frac{2.3 \text{ moles}}{0.45 \text{ L}} = 5.111 \frac{\text{mol}}{\text{L}} = \boxed{5.11 \text{ M}}$$



- 2) 1.2 moles of calcium carbonate in 1.22 liters of solution.

$$M = \frac{1.2 \text{ mol}}{1.22 \text{ L}} = 0.983606557 \frac{\text{mol}}{\text{L}} = \boxed{0.98 \text{ M}}$$



- 3) 0.09 moles of sodium sulfate in 12 mL of solution.

12 mL = 0.012 L

$$M = \frac{0.09 \text{ mol}}{0.012 \text{ L}} = \boxed{7.5 \text{ M}}$$



- 4) 0.75 moles of lithium fluoride in 65 mL of solution.

65 mL = 0.065 L

$$M = \frac{0.75 \text{ mol}}{0.065 \text{ L}} = 11.53846154 \text{ M} = \boxed{11.54 \text{ M}}$$



- 5) 0.8 moles of magnesium acetate in 5 liters of solution.

$$M = \frac{0.8 \text{ mol}}{5 \text{ L}} = \boxed{0.16 \text{ M}}$$



- 6) 120 grams of calcium nitrite in 240 mL of solution.

240 mL = 0.240 L

$$n = 120 \text{ g} \left(\frac{1 \text{ mol}}{132.09 \text{ g}} \right)$$

MW_{Ca(NO₂)₂} = 132.09 g/mol

n = 0.91 mol

$$M = \frac{0.91 \text{ mol}}{0.240 \text{ L}} = \boxed{3.79 \text{ M}}$$



- 7) 98 grams of sodium hydroxide in 2.2 liters of solution.

MW_{NaOH} = 39.997 g/mol

$$n = 98 \text{ g} \left(\frac{1 \text{ mol}}{39.997 \text{ g}} \right) = 2.45 \text{ mol}$$

$$M = \frac{2.45 \text{ mol}}{2.2 \text{ L}} = \boxed{1.11 \text{ M}}$$



- 8) 1.2 grams of hydrochloric acid in 25 mL of solution.

MW_{HCl} = 36.461 g/mol

$$n = (1.2 \text{ g}) \left(\frac{1 \text{ mol}}{36.461 \text{ g}} \right) = 0.033 \text{ mol}$$

$$M = \frac{0.033 \text{ mol}}{0.025 \text{ L}} = \boxed{1.32 \text{ M}}$$



- 9) 45 grams of ammonia in 0.75 L of solution.

MW_{NH₃} = 17.031 g/mol

$$n = 45 \text{ g} \left(\frac{1 \text{ mol}}{17.031 \text{ g}} \right) = 2.64 \text{ mol}$$

$$M = \frac{2.64 \text{ mol}}{0.75 \text{ L}} = \boxed{3.52 \text{ M}}$$

Molarity Calculations Chemfiesta

David B. Cook



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Easy as Measures of Concentration Janelle McAlpine, 2012 Does the concept of a mole escape you Do you have a problem remembering the difference between molarity osmolarity and tonicity Why are they even relevant And what are hypotonic hypertonic and isotonic How are they measured in the body These concepts explained in basic step by step plain English Too easy

An Introduction to Molarity, 2016 One important idea to master in any introductory chemistry course is the concept of concentration of a solution Here you explore this concept the components that make up a solution and learn about a basic unit of measurement for concentration molarity

Methods of Calculation of Molar and Partial Molar Properties from Data for One Component and from Condensed Phase Equilibria N. A. Gokcen, AEROSPACE CORP EL SEGUNDO CALIF LAB OPERATIONS., 1971 Equations have been derived for calculating the excess thermodynamic properties of multicomponent systems either from the data on one component or from the heterogeneous condensed phase equilibria This procedure eliminates the need for existing methods of cumbersome graphical integration for the Gibbs Duhem equations

Author **24 Lessons that Rocked the World** Ian Guch, 1999

Ab Initio Calculations Petr Čárský, Miroslav Urban, 1980

Until recently quantum chemical ab initio calculations were restricted to atoms and very small molecules As late as in 1960 Allen I and Karo stated Almost all of our ab initio experience derives from diatomic LCAO calculations N and we have found in the literature approximately eighty calculations three fourths of which are for diatomic molecules There are approximately twenty ab initio calculations for molecules with more than two atoms but there is a decided dividing line between the existing diatomic and polyatomic wave functions Confidence in the satisfactory evaluation of the many center two electron integrals is very much less than for the diatomic case Among the noted twenty calculations SiH was the largest 4 molecule treated In most cases a minimal basis set was used and the many center two electron integrals were calculated in an approximate way Under these circumstances the ab initio calculations could hardly provide useful chemical information It is therefore no wonder that the dominating role in the field of chemical applications was played by semiempirical and empirical methods The situation changed essentially in the next decade The problem of many center integrals was solved efficient and sophisticated computer programs were developed basis sets suitable for a given type of problem were suggested and meanwhile a considerable amount of results has been accumulated which serve as a valuable comparative material The progress was of course inseparable from the development and availability of computers

Equations for Molar Quantities in Multicomponent Mixtures and Solutions Murugesapillai Maheswaran, 2020

Density-Weight Percent Molarity Conversion Equations for Uranyl Sulfate- Water Solutions at 25.0 C and Between 100-300 C., 1952

Ab Initio Valence Calculations in Chemistry David B. Cook, 1974-01-01

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