

Extension Questions

15. One mole of any gas will occupy 22.4 L of volume at standard temperature and pressure (STP).

Consider this situation as it relates to the reaction in Model 1: $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightarrow 2\text{NH}_3(\text{g})$

a. Calculate the volumes of reactants consumed and the volume of product made.

b. Record the ratio of N_2 to H_2 to NH_3 . Reduce the ratio to the lowest whole numbers possible.

	N_2 Consumed	H_2 Consumed	NH_3 Produced	Volume Ratio $\text{N}_2:\text{H}_2:\text{NH}_3$
How many liters of each substance would be consumed or produced if the reaction occurred 6.02×10^{23} times at STP?	22.4 L	67.2 L	44.8 L	1 : 3 : 2

16. Refer to the data table in Question 15.

a. Can the mole ratio from a balanced chemical equation be interpreted as a ratio of volumes for gases?

Yes. In this case, the mole ratio matches with volume.

b. Use mathematical concepts to explain how your answer in part a is possible.

We are multiplying by the same number (6.02×10^{23}), so we can divide by 22.4 to get back to the ratio of 1 : 3 : 2.

17. Explain why the ratio of volumes is NOT followed in the following reactions.



44.8 L 22.4 L 0.036 L



22.4 L 22.4 L 0.035 L

The volume must be volumes of a gas, and these two reactions produce a liquid and a solid.

18. Which of the following quantities are **conserved** (total amount in reactants = total amount in products) in a chemical reaction? Find an example or counter-example from this activity to support your answer for each.

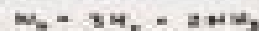
a. Molecules - **does not conserve**

$$1 : 3 : 2 \quad 1 + 3 \neq 2$$

c. Mass - **conserves**

$$19 + 5 = 19 \quad 19 + 5 = 19$$

e. Atoms of an element - **conserves**



$$2\text{N} + 6\text{H} \rightarrow 2\text{N} + 6\text{H}$$

b. Moles - **does not conserve**

$$1 : 3 : 2 \quad 1 + 3 \neq 2$$

d. Volume - **does not conserve**

$$1 : 1 : 2 \quad 1 + 1 \neq 2$$

$$(22.4 + 22.4 \neq 44.8)$$

Mole Ratio Pogil Key

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Mole Ratios Lifeliqe, 2019 This lesson plan relates balanced chemical equations to everyday analogies defines stoichiometry and uses mole ratios to convert between amounts of substances in chemical reactions

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