

10. Which of the following combinations of solutions will produce a precipitate?

- (a) NH_4I and AgNO_3
- (b) KNO_3 and Na_2CO_3
- (c) $\text{Cu}(\text{NO}_3)_2$ and K_2SO_4
- (d) $\text{Fe}(\text{NO}_3)_3$ and MgCl_2

a only

a & c

a, c & d

b, c & d

11. Which of the following substances dissociate in aqueous solution to produce two moles of ions per one mole of dissolved solute?

- (a) NaCl
- (b) NaNO_3
- (c) K_2SO_4
- (d) HClO_4

a only

a & b

a, b & d

all of the above

12. The reaction



- (a) is a neutralization reaction.
- (b) is a precipitation reaction.
- (c) yields water.
- (d) involves Cl^- and Na^+ as spectator ions.

b & c

a & c

b, c & d

a, c & d

13. 50 mL of 0.20M NaOH will react completely with

- (a) 50 mL of 0.10M HCl .
- (b) 100 mL of 0.10M HCl .
- (c) 50 mL of 0.10M H_2SO_4 .
- (d) 100 mL of 0.10M H_2SO_4 .

a & c

a & d

b & c

b & d

14. 100 mL of a 0.20M solution of HNO_3

- (a) will conduct electricity better than 100 mL of a 0.20M solution of H_2SO_4 .
- (b) contains 0.020 moles of NO_3^- .
- (c) can be neutralized by the addition of 100 mL of 0.20M KOH .
- (d) is basic.

all of the above

a & b

b, c & d

a & c

Multiple Choice On Types Of Chemical Reaction

Robert R. Redfield



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