

CHEMISTRY 10TH - Multiple Choice Questions

→ CHEMICAL EQUILIBRIUM

CHAPTER# 01

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01. Which one of the following statements is false about dynamic equilibrium?
- (a) It takes place in a closed container
 - (b) Concentration of reactant and products are not changed
 - (c) Rate of the forward reaction is equal to the rate of the reverse reaction
 - (d) ✓ Equilibrium cannot be disturbed by any external stress
02. Consider the following reaction and indicate which of the following best describes equilibrium constant expression K_c . $4\text{NH}_3 + 5\text{O}_2 \rightleftharpoons 4\text{NO} + 6\text{H}_2\text{O}$
- (a) ✓ $K_c = \frac{[\text{NO}]^4 [\text{H}_2\text{O}]^6}{[\text{NH}_3]^4 [\text{O}_2]^5}$
 - (b) $K_c = \frac{[\text{NH}_3]^4 [\text{O}_2]^5}{[\text{NO}]^4 [\text{H}_2\text{O}]^6}$
 - (c) $K_c = \frac{[\text{NH}_3][\text{O}_2]}{[\text{H}_2\text{O}][\text{NO}]}$
 - (d) $K_c = \frac{[4\text{NO}][6\text{H}_2\text{O}]}{[4\text{NH}_3][5\text{O}_2]}$
03. A reaction which is never goes to completion is known as reversible reaction. Reversible reaction is represented by:
- (a) dotted lines
 - (b) single arrow
 - (c) ✓ double arrow
 - (d) double straight line
04. When the magnitude of K_c is small, indicates:
- (a) ✓ Reaction mixture contain most of the reactant
 - (b) Reaction mixture contain most of the product
 - (c) Reaction mixture contain almost equal amount of reactant and product
 - (d) Reaction goes to completion
05. For which system does the equilibrium constant, K_c has units of concentration?
- (a) ✓ $\text{N}_{2(g)} + 3\text{H}_{2(g)} \rightleftharpoons 2\text{NH}_{3(g)}$
 - (b) $\text{N}_{2(g)} + \text{O}_{2(g)} \rightleftharpoons 2\text{NO}_{(g)}$
 - (c) $\text{H}_{2(g)} + \text{I}_{2(g)} \rightleftharpoons 2\text{HI}_{(g)}$
 - (d) $\text{CO}_{2(g)} + \text{H}_{2(g)} \rightleftharpoons \text{CO}_{(g)} + \text{H}_2\text{O}_{(l)}$

Multiple Choice Questions And Answers In Chemical Equilibrium

S Marginson



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Multiple Choice Questions And Answers In Chemical Equilibrium Book Review: Unveiling the Power of Words

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