

Cleaning methods

Manual cleaning methods appropriate for non-immersible, soft and fragile fixtures should be used. External components that come into contact with moist air and do not come into contact with the patient, should be wiped with a clean, non-abrasive cloth and detergent solution or isopropyl alcohol 70%.

Technical Specification

Parameter	Minimum Value	Maximum Value
Operating Frequency	100 Hz	1000 Hz
Operating Voltage	100 V	240 V
Power	10 W	100 W
Weight	100 g	200 g
Dimensions	100 mm	200 mm
Environmental conditions	Temperature	15 °C - 35 °C
	Humidity	20% - 80%
	Atmospheric pressure	860 hPa - 1060 hPa
Regulation	CE mark	CE mark
	RoHS	RoHS
	REACH	REACH
	WEEE	WEEE
CE mark	CE mark	CE mark
	CE mark	CE mark
	CE mark	CE mark
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PERKINS TONOMETER

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Perkins Tonometer Manual

R Bogdan



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 $2\text{NaCl} + \text{F}_2 \rightarrow 2\text{NaF} + \text{Cl}_2$
 $4\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
 $5\text{Pb} + 2\text{HCl} \rightarrow 2\text{H}_2 + \text{PbCl}_2$
 $6\text{AlBr}_3 + 3\text{K}_2\text{SO}_4 \rightarrow 4\text{KBr} + \text{Al}_2\text{SO}_4 + 3\text{S}$
 $4\text{CH}_4 + \text{O}_2 \rightarrow 2\text{CO}_2 + 2\text{H}_2\text{O}$
 $8\text{C} + 3\text{H}_2\text{SO}_4 \rightarrow 2\text{CO}_2 + 2\text{H}_2\text{O}$

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 $1\text{ N}_2 + \text{F}_2 \rightarrow \text{NF}_3$
 $2\text{C}_6\text{H}_{10}\text{O}_2 + \text{CO}_2 + \text{H}_2\text{O} \rightarrow 3\text{HBr} + \text{KHCO}_3 + \text{H}_2\text{O} + \text{KBr} + \text{CO}_2$
 $\text{GaBr} + \text{NaBr} + \text{SNO}_4 \rightarrow 4\text{NaSO}_3 + \text{Ga}_2\text{SO}_3 + 3\text{S} + 5\text{NF}_3 + \text{SNF}_2 + \text{N}_2\text{O}_3$ using the equation from problem 2 above answer the following

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 $3\text{HBr} + \text{KHCO}_3 + \text{H}_2\text{O} + \text{KBr} + \text{CO}_2 \rightarrow 2\text{GaBr}_3 + 3\text{Na}_2\text{SO}_3 + \text{Ga}_2\text{SO}_3 + 3\text{S} + 6\text{NaBr} + 3\text{SNO}_2 + 2\text{NF}_3 + 3\text{SNF}_2 + \text{N}_2\text{O}_3$ using the following equation
 $2\text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow 2\text{H}_2\text{O} + \text{Na}_2\text{SO}_4$

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 $3\text{H}_2\text{O} + 1\text{KBr} + 1\text{CO}_2 \rightarrow 4\text{H}_2\text{O} + 2\text{GaBr}_3 + 3\text{Na}_2\text{SO}_3 + 1\text{Ga}_2\text{SO}_3 + 3\text{H}_2\text{O}$
 $5\text{NaOH} + 3\text{H}_2\text{SO}_4 \rightarrow 2\text{H}_2\text{O} + \text{Na}_2\text{SO}_4$

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 $3\text{H}_2\text{O} + 1\text{KBr} + 1\text{CO}_2 \rightarrow 4\text{H}_2\text{O} + 2\text{GaBr}_3 + 3\text{Na}_2\text{SO}_3 + 1\text{Ga}_2\text{SO}_3 + 3\text{H}_2\text{O}$
 $5\text{NaOH} + 3\text{H}_2\text{SO}_4 \rightarrow 2\text{H}_2\text{O} + \text{Na}_2\text{SO}_4$

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web google classroom balance the following chemical equation $\text{Mg} + \text{OH}^- \rightarrow \text{Mg} + \text{OH}^-$ $\text{HCl} \rightarrow \text{HCl}$ $\text{MgCl}_2 \rightarrow \text{MgCl}_2$ $\text{H}_2 \rightarrow \text{H}_2$ $\text{O} \rightarrow \text{H}_2\text{O}$ note all reactants and products require a coefficient of at least one stuck review related articles videos or use a hint report a problem $7.4.1 \times \times y \theta \pi 8.5.2.0.9.6$

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web a common type of stoichiometric relationship is the mole ratio which relates the amounts in moles of any two substances in a chemical reaction we can write a mole ratio for a pair of substances by looking at the coefficients in front of each species in

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