



Perkin Elmer 9700 Manual

Jörg Epplen, Thomas Lubjuhn



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web solution a titrant volume 0 mL the solution pH is due to the acid ionization of HCl because this is a strong acid the ionization is complete and the hydronium ion molarity is 0.100 M the pH of the solution is then $\text{pH} = -\log 0.100 = 1.000$ pH log 0.100 = 1.000 b titrant volume 12.50 mL

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Volume of NaOH added (mL)	pH of the solution
4	6.4
6	4.6
5	9.5
9	5.9

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web background plot of the pH of a solution against the volume of titrant added is called a titration curve the pH can be measured directly with a pH meter while titrant is added from a buret for the acids used in this experiment the titrant will always be a 0.1 M solution of the strong base NaOH

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web strong acids and strong bases completely ionize in solution resulting in water and a salt for a weak acid there is only partial ionization the titration curve of a strong base strong acid showed a stretched out curve as it started with a slow gradual change in pH as it reached the equivalence point

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