

COVER LETTER FORMAT

Dear Miss Gil,

My associate from the Raffles Marketing Group James Osterlund has informed me of your need for an Operations Manager for your industry specific brands Conway Cosmetics and eHaulpockets.

In this tough economy, I believe my technical and management background with focus in product marketing, crisis management and front-end customer service resolutions would be a perfect fit for your organization. In 2011, I have achieved sizeable productivity increases for Lapelle Electronics, HughesMerch and TabTax LLC. For these accounts, I have managed to turnaround figures despite the unfavorable decline in the tech market. Working within these confines, I have also built effective core teams under the best interest of the company's vision, time and resources. People know me as a motivated and dedicated leader. As my sucessess are a result of my team's combined effort to show consistency under my guidance.

If you believe you can benefit from a highly aggressive and talented professional with a signature reputation for generating results, then I would welcome the opportunity to meet with you.

Thank you for your time and consideration.

Sincerely,

123 Anywhere St., Any City, State, Country 12345

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AN Whitehead



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$$\begin{aligned} & \text{n}_2 + 3 \text{h}_2 \rightarrow 2 \text{nh}_3 \\ & \text{kcl}_3 + \text{o}_2 \rightarrow 2 \text{nacl} + \text{f}_2 \\ & 2 \text{naf} + \text{cl}_2 \rightarrow 2 \text{h}_2 + \text{o}_2 \\ & 2 \text{h}_2\text{o} + \text{pb} + \text{oh}^- \rightarrow 2 \text{hcl} + \text{h}_2\text{o} + 1 \text{pbcl}_2 \\ & \text{albr}_3 + 3 \text{k}_2\text{so}_4 \rightarrow 6 \text{kbr} + \text{al}_2\text{so}_4 + 3 \text{ch}_4 + 2 \text{o}_2 \\ & 1 \text{co}_2 + 2 \text{h}_2\text{o} + \text{c}_3\text{h}_8 + 5 \text{o}_2 \rightarrow 3 \text{co}_2 + 4 \text{h}_2\text{o} + \text{c}_8\text{h}_{18} + 25 \text{o}_2 \\ & 16 \text{co}_2 + 18 \text{h}_2\text{o} \end{aligned}$$

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web may 28 2020 write a balanced molecular equation describing each of the following chemical reactions
 solid calcium carbonate is heated and decomposes to solid calcium oxide and carbon dioxide gas
 gaseous butane C_4H_{10} reacts with diatomic oxygen gas to yield gaseous carbon dioxide and water vapor

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web the coefficient in a balanced equation is an idea the concept of telling the chemist that if the atoms molecules and compounds are balanced there are balanced amounts of the atoms molecules and compounds on the other side in the product

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$$\begin{aligned} & 1 \text{na} + 3 \text{po}_4 + 3 \text{koh} + 3 \text{naoh} + 1 \text{k} + 3 \text{po}_4 + 1 \text{mgf}_2 + 1 \text{li} + 2 \text{co}_3 + 1 \text{mgco}_3 + 2 \text{lif} + 1 \text{p}_4 + 3 \text{o}_2 + 2 \text{p}_2 + \text{o}_3 + 2 \text{rbno}_3 + 1 \text{bef}_2 + 1 \text{be} + \text{no}_3 + 2 \text{rbf} \end{aligned}$$

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 $1\text{ N}_2 + 3\text{ H}_2 \rightarrow 2\text{ NH}_3$
 $2\text{ KClO}_3 \rightarrow 2\text{ KCl} + 3\text{ O}_2$
 $2\text{ NaCl} + \text{F}_2 \rightarrow 2\text{ NaF} + \text{Cl}_2$
 $2\text{ H}_2\text{O} + 2\text{ Fe} \rightarrow \text{Fe}_2\text{O}_3 + 3\text{ H}_2$
 $2\text{ Pb} + 2\text{ HCl} \rightarrow 2\text{ PbCl}_2 + \text{H}_2$
 $2\text{ Al} + 3\text{ Br}_2 \rightarrow 2\text{ AlBr}_3$
 $2\text{ K}_2\text{SO}_4 \rightarrow 4\text{ K}_2\text{O} + 3\text{ SO}_2 + \text{O}_2$
 $2\text{ H}_2\text{O} + 2\text{ Fe} \rightarrow \text{Fe}_2\text{O}_3 + 3\text{ H}_2$
 $2\text{ Pb} + 2\text{ HCl} \rightarrow 2\text{ PbCl}_2 + \text{H}_2$
 $2\text{ Al} + 3\text{ Br}_2 \rightarrow 2\text{ AlBr}_3$
 $2\text{ K}_2\text{SO}_4 \rightarrow 4\text{ K}_2\text{O} + 3\text{ SO}_2 + \text{O}_2$

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 $1\text{ N}_2 + 3\text{ H}_2 \rightarrow 2\text{ NH}_3$
 $2\text{ KClO}_3 \rightarrow 2\text{ KCl} + 3\text{ O}_2$
 $2\text{ NaCl} + \text{F}_2 \rightarrow 2\text{ NaF} + \text{Cl}_2$
 $2\text{ H}_2\text{O} + 2\text{ Fe} \rightarrow \text{Fe}_2\text{O}_3 + 3\text{ H}_2$
 $2\text{ Pb} + 2\text{ HCl} \rightarrow 2\text{ PbCl}_2 + \text{H}_2$
 $2\text{ Al} + 3\text{ Br}_2 \rightarrow 2\text{ AlBr}_3$
 $2\text{ K}_2\text{SO}_4 \rightarrow 4\text{ K}_2\text{O} + 3\text{ SO}_2 + \text{O}_2$

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 $2\text{ NaCl} + \text{F}_2 \rightarrow 2\text{ NaF} + \text{Cl}_2$
 $2\text{ H}_2\text{O} + 2\text{ Fe} \rightarrow \text{Fe}_2\text{O}_3 + 3\text{ H}_2$
 $2\text{ Pb} + 2\text{ HCl} \rightarrow 2\text{ PbCl}_2 + \text{H}_2$
 $2\text{ Al} + 3\text{ Br}_2 \rightarrow 2\text{ AlBr}_3$
 $2\text{ K}_2\text{SO}_4 \rightarrow 4\text{ K}_2\text{O} + 3\text{ SO}_2 + \text{O}_2$

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 $4\text{ Fe} + 3\text{ O}_2 \rightarrow 2\text{ Fe}_2\text{O}_3$
 $2\text{ Fe} + 2\text{ FeBr}_3 + 3\text{ H}_2\text{SO}_4 \rightarrow 2\text{ Fe}_2\text{O}_3 + 3\text{ H}_2\text{O} + 3\text{ SO}_2$
 $2\text{ Fe} + 3\text{ Cl}_2 \rightarrow 2\text{ FeCl}_3$
 $4\text{ Fe} + 3\text{ O}_2 \rightarrow 2\text{ Fe}_2\text{O}_3$
 $2\text{ Fe} + 2\text{ FeBr}_3 + 3\text{ H}_2\text{SO}_4 \rightarrow 2\text{ Fe}_2\text{O}_3 + 3\text{ H}_2\text{O} + 3\text{ SO}_2$

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 $3\text{ Fe} + 2\text{ O}_2 \rightarrow \text{Fe}_3\text{O}_4$
 $2\text{ Fe} + 3\text{ O}_2 \rightarrow \text{Fe}_2\text{O}_3$
 $2\text{ Fe} + 3\text{ O}_2 \rightarrow \text{Fe}_2\text{O}_3$
 $2\text{ Fe} + 3\text{ O}_2 \rightarrow \text{Fe}_2\text{O}_3$
 $2\text{ Fe} + 3\text{ O}_2 \rightarrow \text{Fe}_2\text{O}_3$
 $2\text{ Fe} + 3\text{ O}_2 \rightarrow \text{Fe}_2\text{O}_3$
 $2\text{ Fe} + 3\text{ O}_2 \rightarrow \text{Fe}_2\text{O}_3$
 $2\text{ Fe} + 3\text{ O}_2 \rightarrow \text{Fe}_2\text{O}_3$
 $2\text{ Fe} + 3\text{ O}_2 \rightarrow \text{Fe}_2\text{O}_3$
 $2\text{ Fe} + 3\text{ O}_2 \rightarrow \text{Fe}_2\text{O}_3$

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web mar 14 2023 the key to balancing chemical equations the ultimate goal for balancing chemical equations is to make both sides of the reaction the reactants and the products equal in the number of atoms per element this stems from the universal law of the conservation of mass which states that matter can neither be created nor destroyed

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