

Determination of concentration of a Dextrose solution of unknown strength using an Ostwald Viscometer.

Aim: To determine the concentration of a Dextrose solution of unknown strength using an Ostwald Viscometer and to study the variation of this solution's viscosity as a function of temperature.

Apparatus Required: 1. Ostwald Viscometer, Thermometer
2. Water, Ice water, glass jar
3. Hot plate, standard volumetric flask 25 ml

Chemicals Required: 1. Dextrose, Distilled water

Theory: Viscosity: The internal friction of a liquid which resist the flow of it is known as viscosity.

The force it implies can be measured in terms of the tangent force which is needed to keep the speed of layers constant, this force is called viscous force.

$$\eta = \frac{\pi \mu^4 P L}{8 \pi V L}$$

$$\text{Volume rate law} = \frac{V}{t} = \frac{P \pi \mu^4}{8 \eta L}$$

V = Volume of the liquid

t = time of flow of liquid

η = Viscosity coefficient

P = Hydrostatic pressure

L = Distance travelled by the liquid during time t .

The units of viscosity are called Poise (P). The official SI unit for absolute viscosity is kg/ms or (Pascal second, Pa.s).

Ostwald viscometer is a commonly used viscometer, which consists of a U-shaped glass tube held vertically. For more accurate

Msc Chemistry Practical Manual

Victor M. Corman



Msc Chemistry Practical Manual:

Practical Physical Chemistry Dr. M. Satish Kumar, Physical Chemistry deals with the relations between the physical properties of substances and their composition The present book is intended to serve as a practical manual for undergraduate and post graduate students I have attempted to assemble the list of experiments from my experience and also have drawn upon the experience of the students who have undergone these laboratory courses and felt the inadequacy of the existing syllabus I am aware that I have not yet exhausted all the experiments that they wanted to place in this book but I had to make a selection keeping the size in consideration This manual is largely structured around the standard experiments of physical chemistry Detailed information on instrumentation kinetics experimental methods and data analysis has been covered I will be happier to take all comments and incorporate them in the further editions

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